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Ionic bonding occurs between a cation and an anion while a covalent bond occurs between two nonmetals. In an ionic bond, one electron is taken from another atom and is completely removed. In a covalent bond, the electron is

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shared between two molecules. An ionic bond forms an ionic compound while a covalent bond forms a molecule.

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A type of weak
chemical bond formed
when the slightly
positive hydrogen
atom of a polar
covalent bond in one
molecule is attracted
to the slightly negative
atom of a polar
covalent bond in

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another molecule.

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covalent. 2 atoms in
single covalent bond
share h.... 2 atoms in
double covalent bond
share h.... 2 atoms in
triple covalent bond
share h.... nonmetals
bond with each other
by sharing electrons to

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become.... 2.

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Bonding 5 Chapter 15 -
16 Assignment &
Problem Set 22. What
is an "alloy"? Chapter
16 Problems Single
Covalent Bonds 23.
Describe the difference
between an ionic and a
covalent bond. 24.
Classify the following
compounds as ionic or

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covalent: a. MgCl_2 b.
 Na_2S c. H_2O d. H_2S
25. Explain why neon is
monatomic but
chlorine is diatomic.
26.

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A covalent bond is the same as a ionic bond. The Octet rule only applies to molecules with covalent bonds. A molecule is polar if the shared electrons are not equally shared. A molecule is nonpolar if the shared electrons are equally shared. Methane gas (CH_4) has a nonpolar covalent bond because

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it is a gas. Solution:
Only d) is true. 5.

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Formal Classification of
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of the Elements .

Malcolm L. H. Green. J.
Organomet. Chem.
1995, 500, 127-148.

Introduction. Covalent
molecules are often
described in terms of
the oxidation state
formalism in which a
charge is assigned to
the atom of interest.

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Work Step by Step In a covalent bond, electrons are shared between the two bonded atoms. While there can still be poles in covalent bonds due to differences in electronegativity, the ionic bond involves the

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Supplemental Work
Optional reading is listed at the end of every chapter of the textbook. Additional problems may be found in the textbook and in any other inorganic chemistry textbook, many of which may be found in

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